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in the nucleus do not change during normal chemical reactions. Only the outer electrons move.

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following examples: 4 a. S in  $\text{H}_2\text{SO}_3$  6 b. S in  $\text{MgSO}_4$  2 c. S in  $\text{K}_2\text{S}$  1 d. Cu in  $\text{Cu}_2\text{S}$  6 e. Cr in  $\text{Na}_2\text{CrO}_4$  5 f. N in  $\text{HNO}_3$  4 g. C in  $(\text{HCO}_3)_3$  3 h. N in  $(\text{NH}_4)_2\text{SO}_4$  2 a.

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Answer: (b) Lewis Base. Explanation: R - NH<sub>2</sub> (Amines) behaves as a Lewis base because it is capable of donating a lone pair of electron. We hope the given NCERT MCQ Questions for Class 11 Chemistry Chapter 7 Equilibrium with Answers Pdf free download will help you.

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As shown in Figure 7.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a

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vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide forms. We can describe this reaction with a chemical equation An expression that gives the identities and quantities of the ...

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